

Analytical Chemistry: Introduction (C003976)

Course size	<i>(nominal values; actual values may depend on programme)</i>		
Credits 4.0	Study time 120 h	Contact hrs	37.5 h

Course offerings and teaching methods in academic year 2023-2024

A (semester 1)	Dutch	Gent	lecture seminar
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Lecturers in academic year 2023-2024

Vandenabeele, Peter	WE06	lecturer-in-charge
Kaczmarek, Anna	WE06	co-lecturer

Offered in the following programmes in 2023-2024

	crdts	offering
Bachelor of Science in Chemistry	4	A
Preparatory Course Master of Science in Chemistry (main subject Analytical and Environmental Chemistry)	4	A

Teaching languages

Dutch

Keywords

Chemical equilibrium, titrimetry (volumetry), gravimetry, sampling, data evaluation

Position of the course

The course Analytical Chemistry: Introduction is the first course within the educational curriculum analytical chemistry. The aim and importance of analytical chemistry is discussed. The course aims at a more profound insight into chemical equilibrium and provides the students with the tools for problem-solving in this context. Important basic concepts in analytical chemistry and their relevance are elucidated. The importance of adequate sampling is highlighted and a number of sampling and sample pretreatment approaches are covered. A profound insight into the wet-chemical analysis techniques titrimetry (or volumetry) and gravimetry is aimed at. Finally, tools are provided for data analysis based on statistics.

Contents

- Introduction
 - o Definition and relevance of analytical chemistry
 - o Qualitative and quantitative analysis
 - o Course of a chemical analysis
- Chemical equilibrium
 - o Introduction
 - o Ladder diagram
 - o A general approach for solving equilibrium problems
 - o Buffer solutions
 - o Activity
- Basic concepts and skills in analytical chemistry
 - o Expressing analysis results (units of concentration)
 - o Stoichiometric calculations
 - o Figures of merit: accuracy, precision, sensitivity, detection and quantification limit, selectivity and sensitivity, robustness

- o Selecting an analytical method
- Data evaluation
- o Measurement uncertainty and distribution of measurement results
- o Statistical analysis of analytical data
- Sampling and sample pretreatment
- o The importance of adequate sampling
- o Sampling strategy
- o Simple methods for separation of analyte and interference
- Gravimetric methods
- o Precipitation gravimetry
- o Volatilization gravimetry
- o Particulate gravimetry
- Titrimetric (volumetric) methods
- o Introduction
- o Acid-base titrations
- o Complexation titrations
- o Redox titrations
- Precipitation titrations

Initial competences

Having followed the courses "Chemistry: Structure of Matter" and "Chemistry II: Reactions or having obtained their competences via equivalent courses.

Final competences

- 1 The student has obtained insight into the aim and importance of analytical chemistry.
- 2 The student has obtained a more profound insight into chemical equilibrium and has enhanced skills for solving equilibrium problems.
- 3 The student is aware of the most important aspects of analysis methods (figures of merit) and analysis data (distribution, uncertainty).
- 4 The student has acquired a profound understanding of the basic principles and the capabilities and limitations of titrimetric (volumetric) and gravimetric methods.
- 5 The student is aware of the importance of adequate sampling and sample preparation prior to the analysis.
- 6 The student can use simple statistical methods for data analysis.

Conditions for credit contract

Access to this course unit via a credit contract is determined after successful competences assessment

Conditions for exam contract

This course unit cannot be taken via an exam contract

Teaching methods

Guided self-study, lecture, seminar: coached exercises, online lecture, online seminar, online seminar: coached exercises

Extra information on the teaching methods

The teaching methods will be dynamically adapted to the situation regarding the COVID19 pandemic. Changed teaching methods can be used if necessary in this context.

Learning materials and price

The students can download the relevant chapters of the course notes 'Analytical Chemistry 2.1' (Harvey). These course notes (in English) are freely distributed via chem.libretexts.org. The PowerPoint files used in the lectures will be made accessible via Ufora. Each student disposes of their own computer.

References

- Analytical Chemistry 2.1, D. Harvey, [http://chem.libretexts.org/Textbook_Maps/Analytical_Chemistry_Textbook_Maps/Map%3A_Analytical_Chemistry_2.0_\(Harvey\)](http://chem.libretexts.org/Textbook_Maps/Analytical_Chemistry_Textbook_Maps/Map%3A_Analytical_Chemistry_2.0_(Harvey))
- Analytical Chemistry, 7th edition, G.D. Christian, P.K. Dasgupta and K.A. Schug, Wiley, 2014.
- Quantitative Chemical Analysis, 9th edition, D.C. Harris and C. Lucy, W.H. Freeman, 2015.

Course content-related study coaching

Questions can be posed via email, after lectures or during a personal meeting after making an appointment by e-mail.

Evaluation methods

end-of-term assessment

Examination methods in case of periodic evaluation during the first examination period

Written examination with open questions, written examination with multiple choice questions, open book examination

Examination methods in case of periodic evaluation during the second examination period

Written examination with open questions, written examination with multiple choice questions, open book examination

Examination methods in case of permanent evaluation**Possibilities of retake in case of permanent evaluation**

not applicable

Extra information on the examination methods

Theoretical part: written examination

Problem-solving part: written examination with open book

Calculation of the examination mark

The theoretical part and the problem-solving part each account for 50% of the total score.
Students who are absent without any well-justified reason or who do not participate to both parts of the evaluation, do not pass the exam for this course unit.